



Course: Organic Chemistry I  
Code: CH203 (joint course with Boston University) 18870 (UAM)  
Location: Faculty of Sciences  
Degree: Chemistry  
Fall 2018-Spring 2019  
Type: Compulsory subject  
Number of credits: 6

## COURSE TITLE

ORGANIC CHEMISTRY-I

### 1.1. Course number

CH203 (joint course with Boston University)/ 18870 UAM

### 1.2. Content area

ORGANIC CHEMISTRY

### 1.3. Course type

Compulsory subject for BU students, optional for UAM students

### 1.4. Course level

Bachelor (second cycle)

### 1.5. Year

2<sup>nd</sup>

### 1.6. Semester

First

### 1.7. Credit allotment

6 ECTS credits

### 1.8. Prerequisites

BU students must have taken at least one of the following options:

Option 1: CH101 (General Chemistry 1) and CH102 (General Chemistry 2).

Option 2: CH101 (General Chemistry 1) and CH110 (General and Quantitative Analytical Chemistry).

Option 3: CH111 (Intensive General and Quantitative Analytical Chemistry 1) and CH112 (Intensive General and Quantitative Analytical Chemistry 2).



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UAM students must have taken any of the following courses:

- 16476 (Química -Grado en Ciencias Ambientales-)
- 16535 (Química -Grado en Ingeniería Química-)
- 16575 (Química -Grado en Ciencias de la Alimentación-)
- 18201 (Química -Grado en Bioquímica-)
- 18422 (Química General -Grado en Nutrición, Biología o Física-)

## 1.9. Minimum attendance requirement

Attendance to classes is strongly recommended. Absences which may result from illnesses, religious holidays, serious accidents, etc must be immediately notified to the Faculty coordinator. To receive full credit a student must document the reason for their absence (for example a signed and stamped doctor's note). Please note that having scheduled exams in another subject is not a valid excuse.

## 1.10. Faculty data

Faculty: M.Victoria Martínez-Díaz (Coordinator).  
Department of Organic Chemistry, Faculty of Sciences, Office: Module 01-610  
Phone: +34 914972436  
Email: victoria.martinez@uam.es  
Website:  
<https://moodle.uam.es/course/search.php?search=organic+chemistry+ch203>  
Office hours: 15-18h appointment previously requested by email

## 1.11. Course objectives

(Taken from BU's Organic Chemistry I Syllabus)

Organic chemistry studies the properties of the compounds of carbon, which alone of all the chemical elements can form the numerous, strong, and directional bonds essential to the construction of biomolecules such as fats, carbohydrates, proteins and nucleic acids. This course focuses on gaining an understanding of the fundamental concepts of the science. Important learning goals that the student will wish to achieve are: (1) familiarity with classical and quantum mechanical theories of bonding; (2) how to name organic compounds systematically; (3) the ability to deduce the structures of organic molecules by interpreting their nuclear magnetic resonance spectra; (4) developing an appreciation for the interplay of an organic molecule's three-dimensional structure and that molecule's chemical properties; (5) the ability to apply



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concepts of physical organic chemistry (e.g., thermodynamics and kinetics) to predict the properties and reactivity of organic compounds. A few reactions (acid-base, radical halogenation, addition to carbon-carbon multiple bonds, nucleophilic substitution and elimination) are covered. The relevance of organic chemistry to biological systems, medicine, environmental science and industry is discussed. In the laboratory portion of the course, students gain experience in conducting organic reactions, purifying products, and interpreting and reporting organic chemical phenomena.

## 1.12. Course contents

### Lecture contents and schedule

Week (Date*)	Topic's name (entry)	Related info
1	Structure and Bonding in Organic Compounds (1)	1.1. Electronic structure. 1.2. Ionic and covalent bonding. Lewis theory and diagrams. Formal charges. 1.3. Inductive effects. Dipole moment. Polar covalent bonds. 1.4. Hybridization: $\sigma$ and $\pi$ bonding 1.5. Drawing organic molecules 1.6. Delocalization of $\pi$ electrons. Resonance. 1.7. Resonance structures. Drawing resonance forms. 1.8. Non-equivalent resonance structures. Rules. 1.9. Aromaticity (MO theory). <i>Chapter 1 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>



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2	Polar Covalent Bonds: Acids and Bases (2)	2.1. Bronsted definition of acids and bases 2.2. Lewis definition of acids and bases 2.3. Organic acids and organic bases 2.4. Acid and base strength 2.5. Nucleophiles and electrophiles <i>Chapter 4 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>
3 - 4	Alkanes and Cycloalkanes (3)	3.1. Alkanes: Nomenclature. Structural isomers. 4.2. Alkanes: Chirality. Optical activity. Cahn-Ingold-Prelog rules. Enantiomers and diastereoisomers. Meso compounds. Conformational isomers. 4.3. Cycloalkanes: Nomenclature. <i>Cis/trans</i> isomerism. 4.4. Cycloalkanes: Stability and ring strain. Conformational isomers. 4.5. Representative examples and physical properties of alkanes and cycloalkanes. 4.6. Reactivity of alkanes: halogenation reaction ( <i>a radical reaction</i> ) <i>Chapters 2 and 3 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>
4 - 5	Structural determination (Nuclear Magnetic Resonance Spectroscopy) (4)	4.1. NMR: principles and spectral features (intensity, chemical shift and couplings). 4.2. Proton and carbon spectra. <i>Chapter 13 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>
6	Haloalkanes (5)	5.1. Nomenclature. Representative examples. Physical properties. Polar covalent bonding. 5.2. Nucleophilic substitution reactions 5.3. Elimination reactions <i>Chapters 8 and 9 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>
7	Alcohols and epoxides (6)	6.1. Nomenclature. Representative examples. Physical properties (acidity and basicity revisited). 6.2. Reactivity: dehydration, oxidation, OH-protection, epoxide ring-opening. <i>Chapters 10 and 11 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>



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8 -10	Alkenes (7)	7.1. Nomenclature. <i>Cis/trans</i> isomerism. 7.2. Representative examples. Physical properties. 7.3. Relative stability. Catalytic Hydrogenation. 7.4. Electrophilic addition reactions: 7.4.1. Mechanism; carbocation stability; regiochemistry; stereochemistry. 7.4.2. Addition of hydrogen halides, halogens and water. 7.4.3. Hydroboration, epoxidation, dihydroxylation, ozonolysis 7.5. Radical additions. <i>Chapters 5 and 6 "Organic Chemistry" Brown, Foote, Iverson,        Anslyn</i>
11 and 13	Alkynes (8)	8.1. Nomenclature. Representative examples. Physical properties. 8.2. Reactivity: catalytic hydrogenation; hydration. 8.3. Alkyne acidity: acetylide anions. 8.4. Alkylation of acetylide anions <i>Chapter 7 "Organic Chemistry" Brown, Foote, Iverson, Anslyn</i>
12	Laboratory	See below
14	Final remarks and exam	Discussions and final exam

(\*) For Fall 2018.

Laboratory contents and schedule:

Date* (Week 12)	Experiment	Related info
12 (11/26- 11/30)	Introduction	Safety and Preparation
	Using Extraction to Isolate the Analgesic Active Principles of a Sample	Acid-base extraction, distillation, recrystallization, melting point, TLC, NMR.
	Chromatography as a Method for Drug Separation	TLC, NMR.
	Separation of the Enantiomers of Ibuprofen	Reflux, vacuum filtration, recrystallization, polarimetry.
	Lab exam (Make-up lab)	



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(\*) For Fall 2018. Specified experiments are tentative and their schedule along the week will be flexible.

### 1.13. Course bibliography

- Organic Chemistry 6th Edition by W. H. Brown, C. S. Foote, B. L. Iverson, E. V. Anslyn (Brooks Cole).
- Organic Chemistry 2nd Edition, by David Klein (Wiley)
- Introduction to Organic Laboratory Techniques. A Microscale Approach, 3<sup>rd</sup> Ed. – BU custom edition, W.B. Saunders, 2000.
- Molecular Modeling Computer Programs.
- e-Learning UAM (Moodle)

## 2. Teaching methodology

1. **Theoretical Lectures:** 50 sessions (5 per week).
2. **Problems Practice:** 20 h (2 h per week). Considering students may have different background levels the number of hours dedicated to practice in-class and homework problems is unusually high.
3. **Laboratory practice:** 5 sessions, 1 week (5 h per session).
4. **Personal Tutorials:** recommended at least 3 per student (ca. 50 min per tutorial)

## 3. Student workload

		N. of hours	Percentage
Compulsory attendance	Theoretical Lectures:	50 h	105 h (70%)
	Problems Practice:	20 h	
	Other: Laboratory practice Mid-Term exams	25 h 6 h	
	Final exam	4 h	



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Personal work	Weekly study and Exam preparation	45 h	45 h (30%)
Total student workload: 25 h x 6 ECTS		150 h	

#### 4. Evaluation procedures and weight of components in the final grade

##### Lecture exams (adapted from BU's Organic Chemistry I Syllabus)

Three 80-minute exams are administered on **M 1 Oct**, **M 29 Oct** and **M 3 Dec** (schedule and location to be announced at least two weeks in advance) in additional time to the lecture meeting time. All lecture exams are cumulative. Lecture exams are graded on a 10-point basis. Your lowest lecture exam score will be dropped in calculating your course grade. No make-up lecture exams are given for any reason: please do not ask to take a make-up lecture exam nor ask to take a lecture exam at other than the scheduled date and time. Do not make travel plans that conflict with the lecture exams.

##### Final exam (adapted from BU's Organic Chemistry I Syllabus to Spanish Law)

A cumulative 4-hour final exam is administered on the morning of **F 14 Dec** (exact time and location to be announced at least two weeks in advance). Final exam is graded on a 10-point basis. See the "Incompletes" section of this syllabus for policies concerning missed final exams.

##### Make-up exams (taken from Dresden w/BU's Organic Chemistry I Syllabus)

There are no make-up exams whatsoever.

##### Lab grades (adapted from Dresden w/BU's Organic Chemistry I Syllabus to Spanish Law)

The grade for the student performance at the laboratory is based on pre-lab questions, experimental skill and ability to record observations about the experiments and post-lab reports to be submitted for each of the assigned experiments. Part of your skill grade will be based on results (e.g. ability to obtain the product) and part will be based on your technique (adherence to safety rules, cleanliness and organizational skills) as assessed by your teaching fellows. Sometimes quizzes will be given at the start of lab and will count as part of the pre-lab preparation grade. Quizzes will be announced on the lab



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information sheet. The grade for the end-of-semester lab exam will count as much as one lab.

### **Make-up Labs** (adapted from Dresden w/BU's Organic Chemistry I Syllabus)

Students are expected to perform all experiments. There are no make-up labs whatsoever in this abroad program since you are leaving the country right after the lab ends. Absences which may result from illnesses, serious accidents, etc must be immediately notified to the Faculty coordinator. To receive full credit a student must document the reason for their absence (for example a signed and stamped doctor's note). Please note having scheduled exams in another subject is not a valid excuse. A student who does not have a documented excuse will receive a maximum grade of 7.0 for the lab absence.

### **Grades** (adapted from BU's Organic Chemistry I Syllabus to Spanish Law)

Course grades are calculated by the following formula:

$$\text{CH203 Grade} = (\text{LE}_1 + \text{LE}_2 + \text{FE} + \text{Lab}) / 4$$

where  $\text{LE}_1$  and  $\text{LE}_2$  are the scores of your two highest of three lecture exams, FE is the score of your final exam, and Lab is your lab score. All individual scores are based on 10 points. The maximum score is 10. The minimum score to pass is 5.0

Spanish grade ranges are:

- Lowest (no pass): 0-4,9: Suspenso (SS).
- Low (pass): 5,0-6,9: Aprobado (AP).
- Intermediate: 7,0-8,9: Notable (NT).
- High: 9,0-9,9: Sobresaliente (SB).
- Highest: 10: Matrícula de Honor (MH)

There are no extra-credit projects to offset poor performance on exams. Please do not ask to have the course grade you earn raised gratuitously because you fail to satisfy the GPA requirements of your program of study, scholarship, etc.

### **Incompletes** (taken from Grenoble w/BU's Organic Chemistry I Syllabus)

Incomplete grades are not accepted in this abroad program, as you will be leaving the site at the end of the semester.





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## 5. Other policies

**Laboratory regulations** (adapted from Dresden w/BU's Organic Chemistry I Syllabus)

**Required Equipment:** To reduce the risk of eye injury approved goggles or safety glasses with side shields must be worn **AT ALL TIMES** in lab. Students ignoring this rule will receive a 1-point reduction in their grade and then will be asked to leave the laboratory if the infraction continues. Contact lenses may **never** be worn in the lab due to safety hazards. Students should own their own safety glasses. A lab coat is required. Closed-toe shoes, long pants, and long-sleeved shirts should be worn into the lab. Your shirt and pants, or a lab coat, must cover you completely from neckline to shoes. Available safety information **MUST** be read prior to the first lab session. **All work should be recorded in a laboratory notebook with consecutively numbered pages. The notebooks will be requested to be turned in at the end of the last session and will not be returned to the students, who are the only responsible for having copies of all work submitted.** Appropriate notebooks will be provided by the laboratory staff.

**Collaboration:** Students are allowed and encouraged to talk about the laboratories together and to collaborate on understanding the material. **However, all work in preparing the pre-labs, observations and reports should be done individually.** No joint preparation of laboratory write-ups is allowed. Any student discovered copying from another student's lab, either past or present, will be taken before the Faculty as will the student whose lab was being copied. Do not "help" other students by handing them your completed lab. You will be held responsible if they copy it.

**Attendance:** Students must always attend the scheduled lab session. Punctuality is crucial as special instructions may be given at the start of lab. Attendance will be taken at the start of the lab and short quizzes may be given. Lab is an integral part of the course. Anyone not completing the laboratory portion of the course will receive an incomplete in the course (see make-up lab section of this syllabus).